1: Which of the following elements is monatomic at room temperature and pressure ?

A: Argon.
B: Chlorine.
C: Hydrogen.
D: Nitrogen
E: Oxygen.

2: The elements in the periodic table are arranged in order of increasing ;

A: Density.
B: Isotope number.
C: Atomic mass.
D: Neutron number.
E: Atomic number.

3: What are the forces of attraction holding helium atoms together called?

A: Hydrogen bonds.
B: London dispersion forces.
C: Ionic interaction.
D: Covalent bonds.
E: Metallic bonds.

4: Which of the following elements consists of discrete covalent molecules ?

A: Lithium
B: Magnesium.
C: Hydrogen.
D: Carbon.
E: Silicon.

5: Which of the following molecules consists of 2 atoms held together by a triple covalent bond ?

A: Oxygen.
B: Nitrogen.
C: Hydrogen.
D: Chlorine.
E: Sulphur.

6: Which of the following molecules consists of polyatomic X_4 molecules?

A: Nitrogen.
B: Carbon.
C: Sulphur.
D: Magnesium.
E: Phosphorus.

7: Which of the following statements explains why diamond has such high melting and boiling points ?

A: Diamond molecules are held together by strong London dispersion forces.
B: The metallic bonds in diamond are very strong.
C: The carbon atoms in diamond are held together by strong covalent bonds.
D: The carbon atoms in diamond are held together by strong hydrogen bonds.
E: The carbon atoms have strong triple covalent bonds holding them together.

8: The layers of carbon atoms in graphite are held together by ;

A: Metallic bonds.
B: Covalent bonds.
C: Hydrogen bonds.
D: London dispersion forces.
E: Electrostatic attraction between ions of opposite charge.